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the difference between a homogeneous mixture and a heterogeneous mixture? 2. List 2 examples of a homogeneous mixture and 2 examples of heterogeneous mixtures. 3. List 2 examples of a pure substance and 2 examples of a mixture. 4.

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Physical Science Final Study Guide. STUDY. PLAY. Kinetic theory. a theory of the behavior of gases that assumes heat is a process of energy transfer and the internal energy of a gas is the total energy of its particles. boiling point. the temperature at which a heated liquid turns into gas.

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a. What is the formula for calculating density? _____

i. What is the density of a cube that is 15 g and has a volume of 3 cm³? _____

ii. Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.

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Physical Science Final Study Guide: Definition. All answers are True: Term. Definition. Metals contain metallic bonds: Term. Definition. Alloys are mixtures of different metals: Term. Definition. In a chemical reaction the mass before = the mass after: Term. Definition. $C + O_2 \rightarrow CO_2$ is a chemical equation:

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Physical Science Final Study Guide 2nd Semester. STUDY. PLAY. Chemical Property. any characteristic of a substance that allows a substance to undergo a change that results in a new substance. To observe some chemical properties, the original substance being observed must _____ into one or more different substances.

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