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Objective questions
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3 - Instantaneous

Rate • We can plot

[C. 4. H. 9. Cl]

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- versus time.
- The rate at any instant in time is called the instantaneous rate.
 - It is the slope of the straight line tangent to the curve at that instant.
 - Instantaneous rate is different from average rate.

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14.4. Use the equations in the AP Chemistry test booklet to work kinetics problems. Explain the concept of reaction half-life and describe the relationship between half-life and rate constant for a first-order reaction. Use graphical analysis

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to determine
whether the rate
law for a reaction is
first or second
order. 14.5-14.7

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Rates • All

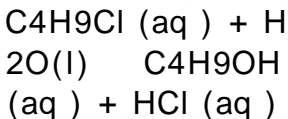
reactions slow

down over time. •

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Therefore, the best indicator of the rate of a reaction is the instantaneous rate near the beginning.



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14.1: Factors that
Affect Reaction

Rates. chemical

kinetics – area of
chemistry dealing

with speeds/rates
of reactions. rates

of reactions

affected by four

factors.

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Concentrations of reactants.
temperature at which reaction occurs. presence of a catalyst. surface area of solid or liquid reactants and/or catalysts.

14.S: Chemical Kinetics

(Summary) -

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Question 1: A catalyst lowers the activation energy of a reaction from 20kJ mol^{-1} to 10kJ mol^{-1} . The temperature at which the catalyzed reaction will have the same rate as that of the catalyzed at $27\text{ }^{\circ}\text{C}$ is $-123\text{ }^{\circ}\text{C}$. 327

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that best completes the statement or answers the question. ... 14)

The rate law for a reaction is $\text{rate} = k[A][B]^2$... The graph shown below depicts the relationship between concentration and time for the following chemical

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reaction. The slope
of this ...

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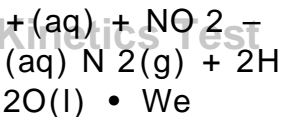
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Consider the

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reaction: NH_4^+



• We measure initial reaction rates. •

The initial rate is the instantaneous rate at time $t = 0$.

• We find this at various initial concentrations of each reactant. • As $[\text{NH}_4^+]$ doubles

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relationship

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... Major topics:

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zero/first/second

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zero/first/second
order reactions
graphically, & half-
life Page 14/22

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It has been found
that for a chemical
reaction with rise in
temperature by 10
° C, the rate

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constant gets nearly doubled. 15. The temperature coefficient of a reaction is the ratio of the rate constants of the reaction at two temperatures differing from one another by 10°C .

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